

12.24 We are asked in this problem to compute the atomic packing factor for the CsCl crystal structure. This requires that we take the ratio of the sphere volume within the unit cell and the total unit cell volume. From Figure 12.3 there is the equivalence of one Cs and one Cl ion per unit cell; the ionic radii of these two ions are 0.170 nm and 0.181 nm, respectively (Table 12.3). Thus, the sphere volume, V_S , is just

$$V_S = \frac{4}{3}(\pi) \left[(0.170 \text{ nm})^3 + (0.181 \text{ nm})^3 \right] = 0.0454 \text{ nm}^3$$

For CsCl the unit cell edge length, a , in terms of the atomic radii is just

$$\begin{aligned} a &= \frac{2r_{\text{Cs}^+} + 2r_{\text{Cl}^-}}{\sqrt{3}} = \frac{2(0.170 \text{ nm}) + 2(0.181 \text{ nm})}{\sqrt{3}} \\ &= 0.405 \text{ nm} \end{aligned}$$

Since $V_C = a^3$

$$V_C = (0.405 \text{ nm})^3 = 0.0664 \text{ nm}^3$$

And, finally the atomic packing factor is just

$$\text{APF} = \frac{V_S}{V_C} = \frac{0.0454 \text{ nm}^3}{0.0664 \text{ nm}^3} = 0.684$$